Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period: \_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_

**Mole Conversions Practice**

| **Conversion** | **Example** | **Convert To:** | **Show Your Work and BOX Your Answer:** |
| --- | --- | --- | --- |
| *1: Moles to Representative Particles* | 5 moles H2O2 | Molecules |  |
| 3 moles Mg | Atoms |  |
| *2: Representative Particles to Moles* | 1.21 x 1023 atoms Ca | Moles |  |
| 3.45 x 1025 formula units XeCl3 | Moles |  |
| *3: Moles to Grams* | 7 moles O2 | Grams | Molar mass O2 = |
| 9 moles Li | Grams | Molar mass Li = |
| *4: Grams to Moles* | 51 grams KCl | Moles | Molar mass KCl = |
| 211 grams Na | Moles | Molar mass Na = |
| *5: Molar Mass* | N2O3 | Grams |  |
| GaBr3 | Grams |  |